

DuPont™ Suva® refrigerants

Thermodynamic Properties of DuPont™ Suva® 404A (HP62) Refrigerant (R-404A)



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SI Units

New tables of the thermodynamic properties of Suva® 404A (HP62) refrigerant [ASHRAE designation: R-404A (44/52/4)], a near azeotropic blend of HFC-125/HFC-143a/HFC-134a, have been developed and are presented here. These tables are based on extensive experimental measurements. Equations have been developed, based on the Peng-Robinson-Stryjek-Vera (PRSV) equation of state, which represent the data with accuracy and consistency throughout the entire range of temperature, pressure, and density presented in these tables.

Physical Properties

Chemical Formula	CHF ₂ CF ₃ /CH ₃ CF ₃ /CH ₂ FCF ₃ (44/52/4% by weight)	
Molecular Weight	97.60	
Boiling Point at One Atmosphere	−46.45°C	(−51.62°F)
Critical Temperature, T _c	72.07°C 345.22 K	(161.73°F) (621.40°R)
Critical Pressure, P _c	3731.5 kPa (abs)	(541.2 psia)
Critical Density, D _c	484.5 kg/m ³	(30.23 lb/ft ³)
Critical Volume, V _c	0.00206 m ³ /kg	(0.0331 ft ³ /lb)

Units and Factors

t = temperature in °C
T = temperature in K = °C + 273.15
p_f = pressure of saturated liquid (bubble point) in kPa (abs)
p_g = pressure of saturated vapor (dew point) in kPa (abs)
v_f = volume of saturated liquid in m³/kg
v_g = volume of saturated vapor in m³/kg
V = volume of superheated vapor in m³/kg
d_f = 1/v_f = density of saturated liquid in kg/m³
d_g = 1/v_g = density of saturated vapor in kg/m³
h_f = enthalpy of saturated liquid in kJ/kg
h_{fg} = enthalpy of vaporization in kJ/kg
h_g = enthalpy of saturated vapor in kJ/kg
H = enthalpy of superheated vapor in kJ/kg
s_f = entropy of saturated liquid in kJ/(kg) (K)
s_g = entropy of saturated vapor in kJ/(kg) (K)
S = entropy of superheated vapor in kJ/(kg) (K)
C_p = heat capacity at constant pressure in kJ/(kg) (K)
C_v = heat capacity at constant volume in kJ/(kg) (K)

The gas constant, R = 8.314 J/(mole) (K)
for Suva® 404A (HP62), R = 0.0852 kJ/kg · K
One atmosphere = 101.325 kPa

Reference point for enthalpy and entropy:

$$h_f = 200 \text{ kJ/kg at } 0^\circ\text{C}$$
$$s_f = 1 \text{ kJ/kg} \cdot \text{K at } 0^\circ\text{C}$$

Equations

The Peng-Robinson-Stryjek-Vera (PRSV) equation of state was used to calculate the tables of thermodynamic properties. It was chosen as the preferred equation of state because it provided an accurate fit of the thermodynamic data over the entire range of temperatures and pressures presented in these tables.

The constants for the PRSV equation of state were calculated in SI units. For conversion of thermodynamic properties to Engineering (I/P) units, conversion factors are provided for each property derived from the PRSV equation of state.

1. Equation of State (PRSV)

$$P = RT/(V - b) - a/(V^2 + 2bV - b^2)$$

where P is in kPa, T is in K, V is in m³/mole, and R = 0.008314 kJ/(mole) (K). The constants a and b are calculated as follows:

$$a = \sum_{i=1}^3 \sum_{j=1}^3 x_i x_j a_{ij} \quad b = \sum_{i=1}^3 x_i b_i$$

where

$$a_{ij} = (a_i a_j)^{0.5} (1 - k_{ij}) \quad b_i = 0.077796 RT_{ci}/P_{ci}$$

x_i = mole fraction of component i

x_j = mole fraction of component j

$$a_i = (0.457235 R^2 T_{ci}^{-2}/P_{ci}) \alpha_i$$

$$a_j = (0.457235 R^2 T_{cj}^{-2}/P_{cj}) \alpha_j$$

k_{ij} = binary interaction parameter for components i and j

$$\alpha_i = [1 + \kappa_i (1 - T_{ri}^{0.5})]^2$$

$$\kappa_i = \kappa_{0i} + \kappa_{1i} [(1 + T_{ri}^{0.5}) (0.7 - T_{ri})]$$

(Note: $\kappa_i = \kappa_{0i}$ for $T_r > 0.7$)

$$\kappa_{0i} = 0.378893 + 1.4897153\omega_i - 0.17131848\omega_i^2 + 0.0196554\omega_i^3$$

κ_{1i} = adjustable parameter for component i

$$T_{ri} = T_i/T_{ci} \text{ for component i}$$

Values for R, T_{ci} , P_{ci} , ω_i , κ_{li} , x_i , and k_{ij} are needed to calculate constants a and b. $R = 0.008314 \text{ kJ}/(\text{mole} \cdot \text{K})$. The remaining constants for Suva® 404A (HP62) are summarized below:

Component	T_{ci}	P_{ci}	ω_i	κ_{li}	x_i
HFC-125 (i = 1)	339.19	3595.0	0.3023	0.0310	0.35782
HFC-143a (i = 2)	346.25	3758.1	0.2529	0.0450	0.60392
HFC-134a (i = 3)	374.20	4056.0	0.3266	-0.0060	0.03826

The binary interaction parameters, k_{ij} , for Suva® 404A (HP62) are:

$$\begin{aligned} k_{11} &= 0.0000 & k_{12} &= -0.0111 & k_{13} &= -0.0024 \\ k_{21} &= -0.0111 & k_{22} &= 0.0000 & k_{23} &= 0.0013 \\ k_{31} &= -0.0024 & k_{32} &= 0.0013 & k_{33} &= 0.0000 \end{aligned}$$

Ideal Gas Heat Capacity Equation (at constant pressure):

$$C_p^o(\text{mixture}) = \sum_{i=1}^3 x_i C_{pi}^o$$

$$C_{pi}^o = 4.184 (A_i + B_i T + C_i T^2 + D_i T^3 + E_i T^4 + F_i T^5)$$

where C_p^o and C_{pi}^o are in J/(mole) (K) and T is in K. x_i is the mole fraction of component i in the mixture [use same values listed in PRSV constants for Suva® 404A (HP62)].

A_i , B_i , C_i , D_i , E_i , and F_i are constants:

$$\begin{aligned} A_1 &= 1.170144 \text{ E+01} & B_1 &= 0.216411 \text{ E-01} \\ A_2 &= 1.372849 \text{ E+00} & B_2 &= 0.750717 \text{ E-01} \\ A_3 &= 4.636855 \text{ E+00} & B_3 &= 0.617904 \text{ E-01} \\ C_1 &= 8.685258 \text{ E-05} & D_1 &= -1.127756 \text{ E-07} \\ C_2 &= -6.206979 \text{ E-05} & D_2 &= 2.011233 \text{ E-08} \\ C_3 &= -3.099070 \text{ E-05} & D_3 &= 0.000000 \text{ E+00} \\ E_1 &= 0.000000 \text{ E+00} & F_1 &= 0.000000 \text{ E+00} \\ E_2 &= 0.000000 \text{ E+00} & F_2 &= 0.000000 \text{ E+00} \\ E_3 &= 0.000000 \text{ E+00} & F_3 &= 0.000000 \text{ E+00} \end{aligned}$$

Properties calculated in SI units from the equations and constants listed above can be converted to I/P units using the conversion factors shown below. Please note that in converting enthalpy and entropy from SI to I/P units, a change in reference states must be included (from $H = 200$ and $S = 1$ at 0°C for SI units to $H = 0$ and $S = 0$ at -40°F for I/P units). In the conversion equations below, H (ref) and S (ref) are the saturated liquid enthalpy and entropy at -40°C . For Suva® 404A (HP62): H (ref) = 145.6 kJ/kg and S (ref) = 0.7862 kJ/kg · K.

Conversion Factors (SI units to I/P units):

P (psia)	$= P$ (kPa) $\cdot 0.14504$
T ($^\circ\text{F}$)	$= (T [^\circ\text{C}] \cdot 1.8) + 32$
D (lb/ft ³)	$= D$ (kg/m ³) $\cdot 0.062428$
V (ft ³ /lb)	$= V$ (m ³ /kg) $\cdot 16.018$
H (Btu/lb)	$= [H (\text{kJ/kg}) - H (\text{ref})] \cdot 0.43021$
S (Btu/lb · $^\circ\text{R}$)	$= [S (\text{kJ/kg} \cdot \text{K}) - S (\text{ref})] \cdot 0.23901$
C_p (Btu/lb · $^\circ\text{F}$)	$= C_p (\text{kJ/kg} \cdot \text{K}) \cdot 0.23901$
C_v (Btu/lb · $^\circ\text{F}$)	$= C_v (\text{kJ/kg} \cdot \text{K}) \cdot 0.23901$

2. Vapor Pressure

$$\log_n P = A + B/T + C \log_n T + D T^2$$

For SI units

T is in K and P is in kPa (abs)

A, B, C and D are constants.

Constants for vapor pressure of saturated liquid (bubble point), p_f :

$$\begin{aligned} A &= 5.56487 \text{ E+01} & C &= -6.58061 \text{ E+00} \\ B &= -3.62385 \text{ E+03} & D &= 1.27711 \text{ E-05} \end{aligned}$$

Constants for vapor pressure of saturated vapor (dew point), p_g :

$$\begin{aligned} A &= 6.89227 \text{ E+01} & C &= -8.71773 \text{ E+00} \\ B &= -4.06171 \text{ E+03} & D &= 1.68264 \text{ E-05} \end{aligned}$$

For I/P units

T is in $^\circ\text{R}$ and P is in psia

A, B, C and D are constants.

Constants for vapor pressure of saturated liquid (bubble point), p_f :

$$\begin{aligned} A &= 5.75859 \text{ E+01} & C &= -6.58061 \text{ E+00} \\ B &= -6.52292 \text{ E+03} & D &= 3.94176 \text{ E-06} \end{aligned}$$

Constants for vapor pressure of saturated vapor (dew point), p_g :

$$\begin{aligned} A &= 7.21161 \text{ E+01} & C &= -8.71773 \text{ E+00} \\ B &= -7.31107 \text{ E+03} & D &= 5.19336 \text{ E-06} \end{aligned}$$

3. Density of the Saturated Liquid

$$d_f/D_c = a_0 + a_1 z + a_2 z^2 + a_3 z^3 + a_4 z^4$$

$$\text{where } z = (1 - T/T_c)^{1/3} - t_0$$

Because both density and temperature appear in the reduced form in the equation, the same constants can be used for either SI or I/P units.

d_f and D_c are in kg/m^3 in SI units and lb/ft^3 in I/P units; T and T_c are in K in SI units and ${}^\circ\text{R}$ in I/P units; a_0, a_1, a_2, a_3, a_4 , and t_0 are constants:

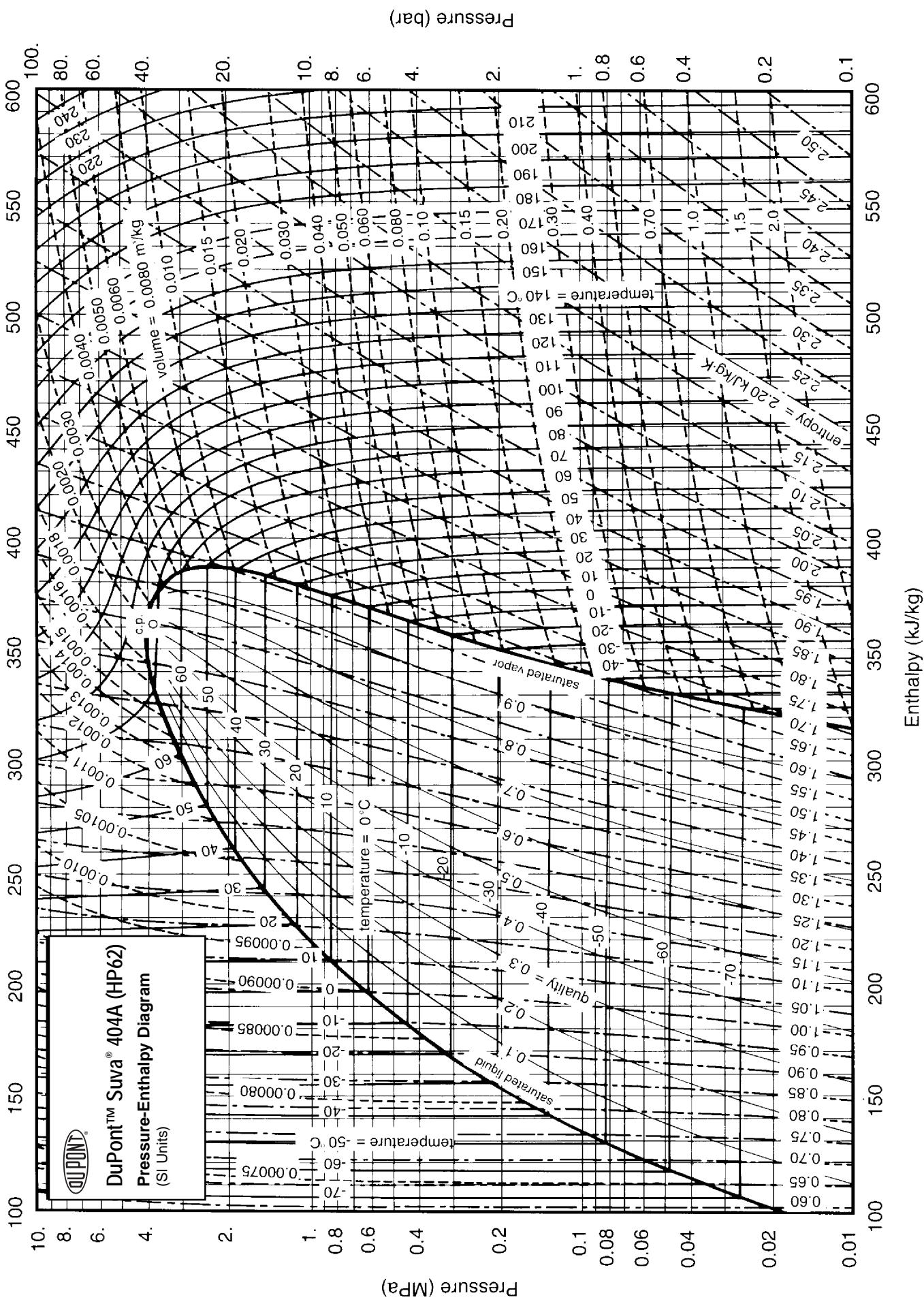
$$a_0 = 1.0002 \quad E+00 \quad a_3 = -1.3781 \quad E+01$$

$$a_1 = 1.9300 \quad E-01 \quad a_4 = 7.6142 \quad E+00$$

$$a_2 = 9.0829 \quad E+00 \quad t_0 = 0.0000$$

Table 1 (continued)
Suva® 404A (HP62) Saturation Properties—Temperature Table

TEMP. °C	PRESSURE kPa		VOLUME m³/kg		DENSITY kg/m³		ENTHALPY kJ/kg			ENTROPY kJ/(kg)(K)		TEMP. °C
	Liquid Pf	Vapor Pg	Liquid Vf	Vapor Vg	Liquid 1/Vf	Vapor 1/Vg	Liquid hf	Latent hfg	Vapor hg	Liquid sf	Vapor sg	
20	1097.7	1085.1	0.0009	0.0181	1071.7	55.267	229.9	148.4	378.3	1.1038	1.6106	20
21	1127.9	1115.2	0.0009	0.0176	1067.2	56.930	231.5	147.3	378.7	1.1091	1.6103	21
22	1158.7	1145.9	0.0009	0.0171	1062.6	58.640	233.0	146.1	379.2	1.1144	1.6099	22
23	1190.2	1177.2	0.0009	0.0166	1057.9	60.398	234.6	144.9	379.6	1.1197	1.6095	23
24	1222.3	1209.2	0.0009	0.0161	1053.2	62.206	236.3	143.7	380.0	1.1250	1.6091	24
25	1255.0	1241.8	0.0010	0.0156	1048.4	64.066	237.9	142.5	380.4	1.1304	1.6087	25
26	1288.4	1275.1	0.0010	0.0152	1043.5	65.980	239.5	141.3	380.8	1.1357	1.6083	26
27	1322.5	1309.1	0.0010	0.0147	1038.5	67.949	241.2	140.0	381.1	1.1411	1.6079	27
28	1357.2	1343.7	0.0010	0.0143	1033.5	69.975	242.8	138.7	381.5	1.1465	1.6075	28
29	1392.6	1379.0	0.0010	0.0139	1028.3	72.062	244.5	137.4	381.9	1.1519	1.6070	29
30	1428.7	1415.0	0.0010	0.0135	1023.1	74.210	246.2	136.1	382.2	1.1574	1.6065	30
31	1465.4	1451.7	0.0010	0.0131	1017.8	76.422	247.9	134.7	382.6	1.1628	1.6060	31
32	1502.9	1489.1	0.0010	0.0127	1012.3	78.702	249.6	133.3	382.9	1.1683	1.6055	32
33	1541.1	1527.2	0.0010	0.0123	1006.8	81.050	251.3	131.9	383.2	1.1738	1.6050	33
34	1580.0	1566.0	0.0010	0.0120	1001.1	83.472	253.0	130.5	383.5	1.1793	1.6044	34
35	1619.7	1605.6	0.0010	0.0116	995.4	85.968	254.8	129.0	383.8	1.1848	1.6038	35
36	1660.1	1645.9	0.0010	0.0113	989.5	88.543	256.5	127.5	384.1	1.1904	1.6032	36
37	1701.2	1687.0	0.0010	0.0110	983.5	91.201	258.3	126.0	384.3	1.1960	1.6026	37
38	1743.1	1728.8	0.0010	0.0106	977.4	93.944	260.1	124.5	384.6	1.2016	1.6019	38
39	1785.8	1771.4	0.0010	0.0103	971.1	96.777	261.9	122.9	384.8	1.2073	1.6012	39
40	1829.2	1814.8	0.0010	0.0100	964.7	99.704	263.8	121.3	385.0	1.2130	1.6005	40
41	1873.4	1859.0	0.0010	0.0097	958.2	102.730	265.6	119.6	385.2	1.2187	1.5998	41
42	1918.4	1904.0	0.0011	0.0094	951.5	105.860	267.5	117.9	385.4	1.2245	1.5990	42
43	1964.2	1949.8	0.0011	0.0092	944.6	109.098	269.4	116.2	385.6	1.2303	1.5981	43
44	2010.8	1996.4	0.0011	0.0089	937.6	112.452	271.3	114.4	385.7	1.2362	1.5973	44
45	2058.3	2043.9	0.0011	0.0086	930.4	115.926	273.2	112.6	385.8	1.2421	1.5964	45
46	2106.6	2092.2	0.0011	0.0084	923.0	119.529	275.1	110.8	385.9	1.2480	1.5954	46
47	2155.7	2141.3	0.0011	0.0081	915.5	123.267	277.1	108.9	386.0	1.2540	1.5944	47
48	2205.6	2191.3	0.0011	0.0079	907.7	127.150	279.1	107.0	386.1	1.2600	1.5933	48
49	2256.5	2242.2	0.0011	0.0076	899.7	131.185	281.1	105.0	386.1	1.2662	1.5922	49
50	2308.2	2294.0	0.0011	0.0074	891.5	135.384	283.2	102.9	386.1	1.2723	1.5910	50
51	2360.7	2346.6	0.0011	0.0072	883.0	139.757	285.3	100.8	386.1	1.2786	1.5897	51
52	2414.2	2400.2	0.0011	0.0069	874.3	144.317	287.4	98.6	386.0	1.2849	1.5884	52
53	2468.6	2454.6	0.0012	0.0067	865.3	149.078	289.6	96.4	385.9	1.2913	1.5870	53
54	2523.8	2510.0	0.0012	0.0065	856.0	154.056	291.7	94.1	385.8	1.2977	1.5855	54
55	2580.0	2566.4	0.0012	0.0063	846.4	159.270	294.0	91.7	385.7	1.3043	1.5839	55
56	2637.1	2623.7	0.0012	0.0061	836.5	164.738	296.2	89.2	385.5	1.3110	1.5822	56
57	2695.2	2681.9	0.0012	0.0059	826.2	170.486	298.5	86.7	385.2	1.3178	1.5804	57
58	2754.2	2741.1	0.0012	0.0057	815.6	176.541	300.9	84.0	384.9	1.3247	1.5785	58
59	2814.2	2801.4	0.0012	0.0055	804.5	182.935	303.3	81.2	384.6	1.3317	1.5764	59
60	2875.1	2862.6	0.0013	0.0053	792.9	189.706	305.8	78.3	384.2	1.3389	1.5742	60
61	2937.0	2924.8	0.0013	0.0051	780.9	196.901	308.4	75.3	383.7	1.3463	1.5718	61
62	2999.9	2988.0	0.0013	0.0049	768.3	204.575	311.0	72.1	383.1	1.3539	1.5692	62
63	3063.8	3052.3	0.0013	0.0047	755.0	212.799	313.7	68.8	382.5	1.3617	1.5664	63
64	3128.7	3117.6	0.0013	0.0045	741.0	221.659	316.5	65.2	381.8	1.3697	1.5633	64
65	3194.6	3184.0	0.0014	0.0043	726.2	231.271	319.5	61.5	380.9	1.3781	1.5599	65



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